

*Heat*

1. A small bag containing 0.200 kg of lead shot at a temperature of 15.0°C falls from a 40.0 m high tower. Instead of bouncing back, the bag makes a small hole in the ground. The specific heat of lead is  $1.28 \times 10^2 \text{ J/kg}\cdot^\circ\text{C}$ .

a. Find the initial potential energy of the lead.

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b. How much energy did the lead lose as heat?

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c. The temperature of the lead after impact was 17.0°C. What was the increase in internal energy of the lead? How does it compare to the amount of lost potential energy?

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d. How much internal energy was added to the ground?

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2. A very shallow pond contains  $1.50 \times 10^5 \text{ kg}$  of water at 23°C. At the end of a windy day,  $1.00 \times 10^3 \text{ kg}$  of water was lost by evaporation. It takes  $2.26 \times 10^6 \text{ J}$  for 1 kg of water to evaporate.

a. How much energy was removed from the pond by heat of evaporation?

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b. How much water was left in the pond?

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c. By how much did the temperature of the water drop in the pond?  
(Hint: the specific heat capacity for water is  $4.19 \times 10^3 \text{ J}/(\text{kg}\cdot^\circ\text{C})$ .)

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d. Assuming there were no other changes in energy, what was the temperature of the water at the end of the day?

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Chapter **10**

**HOLT PHYSICS**  
**Mixed Review** *continued*

**3.** Exactly two kilograms of boiling water ( $100.0^{\circ}\text{C}$ ) are poured into a long, insulated aluminum pipe. The mass of the pipe is  $5.000\text{ kg}$ , and its temperature is  $20.0^{\circ}\text{C}$ . The specific heat capacity of water is  $4.19 \times 10^3\text{ J/kg}\cdot^{\circ}\text{C}$ , and the specific heat capacity of aluminum is  $8.99 \times 10^2\text{ J/kg}\cdot^{\circ}\text{C}$ .

**a.** Given that the final temperature of the water is  $x^{\circ}\text{C}$  and the final temperature of the pipe is  $y^{\circ}\text{C}$ , explain why  $y = x$ .

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**b.** Write expressions for the temperature change in water and in the pipe itself.

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**c.** Write an expression for the amount of energy removed from the water.

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**d.** Write an expression for the amount of energy added to the aluminum.

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**e.** Explain under what conditions these two amounts of energy may be considered equal.

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**f.** Assuming that these conditions are realized, find the final temperature of the water and pipe.

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